**Yr 11 VALENCY**

Students should be able to recognise and write the formula of the following ions and molecules:

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| **Ion name** | **Formula** |  | **Ion name** | **Formula** |
| ammonium |  |  | bromide |  |
| caesium |  |  | chloride |  |
| hydrogen |  |  | cyanide |  |
| lithium |  |  | dihydrogenphosphate |  |
| potassium |  |  | ethanoate (acetate) |  |
| rubidium |  |  | fluoride |  |
| silver |  |  | hydrogencarbonate |  |
| sodium |  |  | hydrogensulfate |  |
| barium |  |  | hydroxide |  |
| calcium |  |  | iodide |  |
| cobalt(II) |  |  | nitrate |  |
| copper(II) |  |  | nitrite |  |
| iron(II) |  |  | permanganate |  |
| lead(II) |  |  | carbonate |  |
| magnesium |  |  | chromate |  |
| manganese(II) |  |  | dichromate |  |
| nickel(II) |  |  | hydrogenphosphate |  |
| strontium |  |  | oxalate |  |
| zinc |  |  | oxide |  |
| aluminium |  |  | sulfate |  |
| chromium(III) |  |  | sulfide |  |
| iron(III) |  |  | sulfite |  |
|  |  |  | nitride |  |
|  |  |  | phosphate |  |

**Common molecules that have non-systematic names:**

|  |  |
| --- | --- |
| **Molecule name** | **Formula** |
| ammonia |  |
| water | O |
| hydrogen peroxide |  |
| ethanoic acid |  |
| hydrochloric acid |  |
| nitric acid |  |
| carbonic acid |  |
| sulfuric acid |  |
| sulfurous acid |  |
| phosphoric acid |  |